Review topic: Periodic Table

	Ways to practice skills	R	A	G	Comment
2.2 Atomic structure and the Periodic Table					
6	State that: a. Group VIII noble gases have a full outer shell b. the number of outer shell electrons is equal to the group number in Groups I to VII c. the number of occupied electron shells is equal to the period number	☐	☐	☐	
8.1 Arrangement of elements					
1	Describe the Periodic Table as an arrangement of elements in periods and groups and in order of increasing proton / atomic number	☐	☐	☐	
2	Describe the change from metallic to non-metallic character across a period	☐	☐	☐	
3	Describe the relationship between group number and the charge of the ions formed from elements in that group	☐	☐	☐	
4	Explain similarities in the chemical properties of elements in the same group of the Periodic Table in terms of their electronic configuration	☐	☐	☐	
5	Explain how the position of an element in the Periodic Table can be used to predict its properties	☐	☐	☐	
6	Identify trends in groups, given information about the elements	☐	☐	☐	
8.2 Group I properties					
1	Describe the Group I alkali metals, lithium, sodium and potassium, as relatively soft metals with general trends down the group, limited to: a. decreasing melting point b. increasingdensity c. increasing reactivity	☐	☐	☐	
2	Predict the properties of other elements in Group I, given information about the elements	☐	☐	☐	
8.3 Group VII properties					
1	Describe the Group VII halogens, chlorine, bromine and iodine, as diatomic non- metals with general trends down the group, limited to: a. increasingdensity b. decreasingreactivity	☐	☐	☐	
2	State the appearance of the halogens at r.t.p. as: a. chlorine, a pale yellow-green gas b. bromine, a red-brown liquid c. iodine, a grey-black solid	☐	☐	☐	

	Ways to practice skills	R	A	G	Comment
3	Describe and explain the displacement reactions of halogens with other halide ions	☐	☐	☐	
4	Predict the properties of other elements in Group VII, given information about the elements	☐	☐	☐	
8.4 Transition elements					
1	Describe the transition elements as metals that: a. have high densities b. have high melting points c. form coloured compounds d. often act as catalysts as elements and in compounds	☐	☐	☐	
2	Describe transition elements as having ions with variable oxidation numbers, including iron(II) and iron(III)	☐	☐	☐	
8.5 Noble gases					
1	Describe the Group VIII noble gases as unreactive, monatomic gases and explain this in terms of electronic configuration	☐	☐	☐	